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Level: SL IB in Biology  
Subject: Biology  
Topic: IB SL Biology  
Type: Mark Scheme

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All International Baccalaureate IB Topic Questions SL Biology

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**BIOLOGY**

**SL - IB**

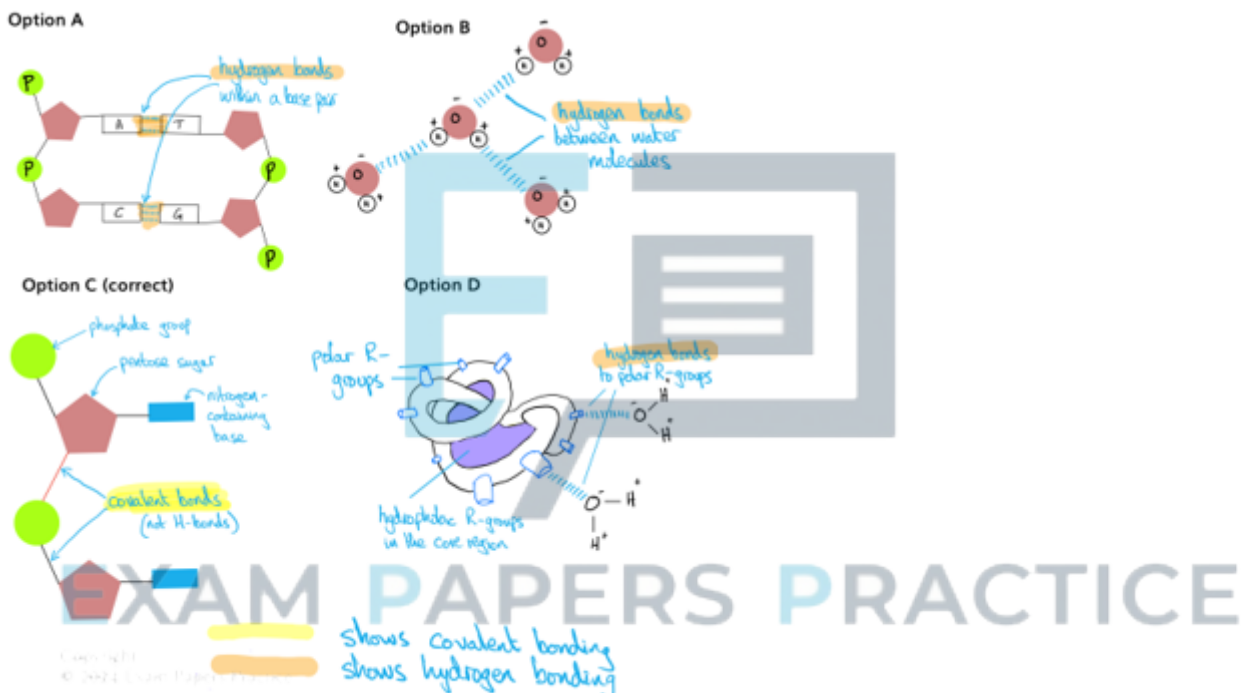
Key skills

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## Mark Scheme

### Answer 1

The correct answer is C because nucleotides are joined to each other in a strand by strong, covalent bonds. The rest are examples of hydrogen bonding.



### Answer 2

The correct answer is B because the lack of polarity means that hydrogen bonds do not form between methane molecules, as they do in water. Therefore very less energy is required to separate molecules from the liquid state to the gaseous state, hence methane's low boiling point (-162°C).

A is incorrect because a lot of compounds have a low molecular weight, and this is not related to their polarity.

C is incorrect because flammability is linked to the energy content of the covalent bonds within the molecule.

D is incorrect because many gases, of varying polarity, contribute to the greenhouse effect and climate change, notably CO<sub>2</sub>, water vapour and methane, which have different levels of polarity.

Answer 3

The correct answer is A because ice crystals form in a way that makes them less densely packed than in liquid water. This is not caused by water's high latent heat of vaporisation or specific heat capacity.

B is incorrect because water's high latent heat of vaporisation causes its boiling point to be 100°C, so at temperatures that exist naturally on Earth, water can form, ice, liquid water or steam (water vapour).

C is incorrect because it is this feature that allows a lot of cooling effect by secreting a small volume of water, in sweating (animals) or transpiration (plants).

D is a definition of water's high specific heat capacity.

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Answer 4

The correct answer is D because sodium chloride (salt) is the most soluble, and cholesterol is insoluble in water. A single molecule of an amino acid is still more soluble than oxygen, as oxygen requires the assistance of haemoglobin to transport to cells and tissues.

Cholesterol is the least soluble - eliminate A

Sodium chloride (salt) is the most soluble - eliminate C

Oxygen is less soluble than all amino acids - eliminate B

Answer 5

The correct answer is D because all the properties are caused by forces that act between molecules of water. These are hydrogen bonds.



Surface tension relies on hydrogen bonds that pull water molecules together and create a 'skin' on the surface. Hydrogen bonds make water a good solvent by bonding with polar solutes, surrounding the individual solute particles. Water has strong cohesive properties because it attracts other water molecules strongly through hydrogen bonds. A large amount of heat is required to break the strong hydrogen bonds in water, explaining water's high specific heat capacity.

#### Answer 6

The correct answer is D because, as water has a high specific heat capacity, a lot of heat is required to raise its temperature. This helps stabilise the temperature that enzymes work in, minimising denaturation by exposure to heat.

(i) is incorrect because water does retain a lot of heat, but this would lead to the temperature in the leaves increasing which could cause enzymes to denature.

(ii) is incorrect as water does not form hydrogen bonds with nonpolar molecules. Even with polar ones, this property of water does not prevent denaturation of enzymes at extremes of temperature or pH.

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(iv) is incorrect because while water molecules are cohesive (they stick together), that doesn't prevent denaturation of enzymes at extremes of temperature or pH.

#### Answer 7:

The correct answer is B. Water's ability to form hydrogen bonds is primarily responsible for its high boiling point compared to other molecules of similar size.

A can be ruled out because water's polarity alone does not account for its high boiling point.

C is incorrect as water's low density is not related to its boiling point.

D is incorrect because the ability to dissolve many substances is not related to its boiling point.



**Answer 8:**

The correct answer is B. Water's high specific heat capacity helps maintain stable internal temperatures by absorbing and releasing heat slowly.

A is incorrect because water does not allow for rapid temperature changes.

C is incorrect as high specific heat capacity does not increase metabolic reaction rates.

D is incorrect because it relates to solubility, not temperature regulation.

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**Answer 9:**

The correct answer is A. Water's cohesion property aids in the process of transpiration in plants.

B is incorrect as cohesion does not directly break down plant cell walls.

C is incorrect because cohesion does not affect the density of plant tissues.

D is incorrect as cohesion does not prevent water droplets on leaves.

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**Answer 10:**

The correct answer is A. Water's hydrogen bonding capability contributes to its ability to dissolve a wide range of substances.

B is incorrect as water's low viscosity does not impact its dissolving power.

C is incorrect because water's high density is not related to its solvent properties.

D is incorrect as water is a polar substance, not non-polar.

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**Answer 11:**

The correct answer is A. Water is known as a universal solvent because it can dissolve both ionic and polar substances.

B is incorrect as water's high boiling point does not make it a universal solvent.

C is incorrect as water's density is unrelated to its solvent properties.

D is incorrect because water remains a liquid at room temperature and this does not relate to its solvent capabilities.



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