

Write your name here

Surname

Other names

Pearson Edexcel Certificate

Centre Number

Candidate Number

**Pearson Edexcel
International GCSE**

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Chemistry

Unit: KCH0/4CH0

Paper: 2C

Wednesday 14 June 2017 – Morning

Time: 1 hour

Paper Reference

**KCH0/2C
4CH0/2C**

You must have:

Calculator

Total Marks

Instructions

- Use **black** ink or ball-point pen.
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided
– *there may be more space than you need.*
- Show all the steps in any calculations and state the units.
- Some questions must be answered with a cross in a box ☒. If you change your mind about an answer, put a line through the box ☒ and then mark your new answer with a cross ☒.

Information

- The total mark for this paper is 60.
- The marks for **each** question are shown in brackets
– *use this as a guide as to how much time to spend on each question.*

Advice

- Read each question carefully before you start to answer it.
- Write your answers neatly and in good English.
- Try to answer every question.
- Check your answers if you have time at the end.

Turn over ►

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THE PERIODIC TABLE

1 2 3 4 5 6 7 8 9 10

Period

Group

1

1
H
Hydrogen
1

4
He
Helium
2

9 Be Beryllium 4	24 Mg Magnesium 12	40 Ca Calcium 20	88 Sr Strontium 38	137 Ba Barium 56	226 Ra Radium 88	227 Ac Actinium 89
7 Li Lithium 3	23 Na Sodium 11	39 K Potassium 19	86 Rb Rubidium 37	133 Cs Caesium 55	223 Fr Francium 87	

11 B Boron 5	12 C Carbon 6	14 N Nitrogen 7	16 O Oxygen 8	19 F Fluorine 9	20 Ne Neon 10
27 Al Aluminium 13	28 Si Silicon 14	31 P Phosphorus 15	32 S Sulfur 16	35.5 Cl Chlorine 17	40 Ar Argon 18
70 Ga Gallium 31	73 Ge Germanium 32	75 As Arsenic 33	79 Se Selenium 34	80 Br Bromine 35	84 Kr Krypton 36
115 In Indium 49	119 Sn Tin 50	122 Sb Antimony 51	128 Te Tellurium 52	127 I Iodine 53	131 Xe Xenon 54
204 Tl Thallium 81	207 Pb Lead 82	209 Bi Bismuth 83	210 Po Polonium 84	210 At Astatine 85	222 Rn Radon 86

65 Zn Zinc 30	63.5 Cu Copper 29	59 Ni Nickel 28	59 Co Cobalt 27	56 Fe Iron 26	55 Mn Manganese 25	52 Cr Chromium 24	51 V Vanadium 23	48 Ti Titanium 22	45 Sc Scandium 21
112 Cd Cadmium 48	108 Ag Silver 47	106 Pd Palladium 46	103 Rh Rhodium 45	101 Ru Ruthenium 44	99 Tc Technetium 43	96 Mo Molybdenum 42	93 Nb Niobium 41	91 Zr Zirconium 40	89 Y Yttrium 39
201 Hg Mercury 80	197 Au Gold 79	195 Pt Platinum 78	192 Ir Iridium 77	190 Os Osmium 76	186 Re Rhenium 75	184 W Tungsten 74	181 Ta Tantalum 73	179 Hf Hafnium 72	139 La Lanthanum 57

Key

Relative atomic mass
Symbol
Name
Atomic number

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Answer ALL questions.

1 Sodium sulfate is a compound with many uses.

(a) The formula of the main compound used as the source of sodium sulfate is $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$

How many different elements are shown in this formula?

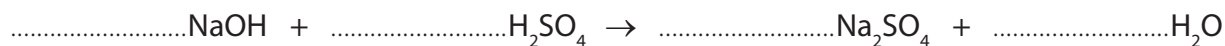
(1)

- A 2
- B 3
- C 4
- D 5

(b) Sodium sulfate can be made from sodium hydroxide and sulfuric acid.

Balance the equation for the reaction between sodium hydroxide and sulfuric acid.

(1)



(c) Sodium hydroxide is manufactured by the electrolysis of brine in the diaphragm cell.

Sulfuric acid is manufactured using the contact process.

The table contains some statements about these two processes.

Place ticks (✓) in the boxes to show the two correct statements.

(2)

brine is a solution of sodium chloride in water	
the temperature used in the contact process is greater than 1000 °C	
an equation for the contact process is $\text{SO}_2 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{SO}_4$	
the reactions in the diaphragm cell are displacement reactions	
the catalyst used in the contact process is vanadium(V) oxide	

(Total for Question 1 = 4 marks)



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2 The diagram shows the positions of some elements in part of the Periodic Table.

							He
Li		B		N		F	
	Mg		Si		S		Ar

(a) How many periods and groups are shown in this diagram?

(1)

	Periods	Groups
<input type="checkbox"/> A	2	4
<input type="checkbox"/> B	3	4
<input type="checkbox"/> C	2	8
<input type="checkbox"/> D	3	8

(b) How many elements shown in the diagram are noble gases?

(1)

- A 1
- B 2
- C 3
- D 4

(c) What is the formula of the compound formed between magnesium and fluorine?

(1)

- A MgF
- B Mg₂F
- C MgF₂
- D Mg₂F₂

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(d) The table shows the percentage composition by mass of a sample of silicon.

Isotope	^{28}Si	^{29}Si	^{30}Si
Percentage (%)	92.2	4.70	3.10

Calculate the relative atomic mass of this sample of silicon.

Give your answer to one decimal place.

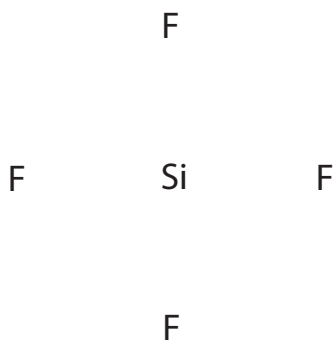
(2)

relative atomic mass =

(e) A molecule of silicon tetrafluoride (SiF_4) contains covalent bonds.

Draw a dot and cross diagram to show the outer electrons in this molecule.

(2)



- (f) The table shows the boiling points of some compounds containing silicon. All of these compounds contain covalent bonds.

Compound	Boiling point in °C
SiF_4	-86
SiCl_4	58
SiO_2	2950

SiF_4 and SiCl_4 have simple molecular structures.

SiO_2 has a giant covalent structure.

- (i) Explain why the boiling point of SiCl_4 is greater than the boiling point of SiF_4 (2)

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- (ii) Explain why the boiling point of SiO_2 is very much greater than the boiling point of SiCl_4 (2)

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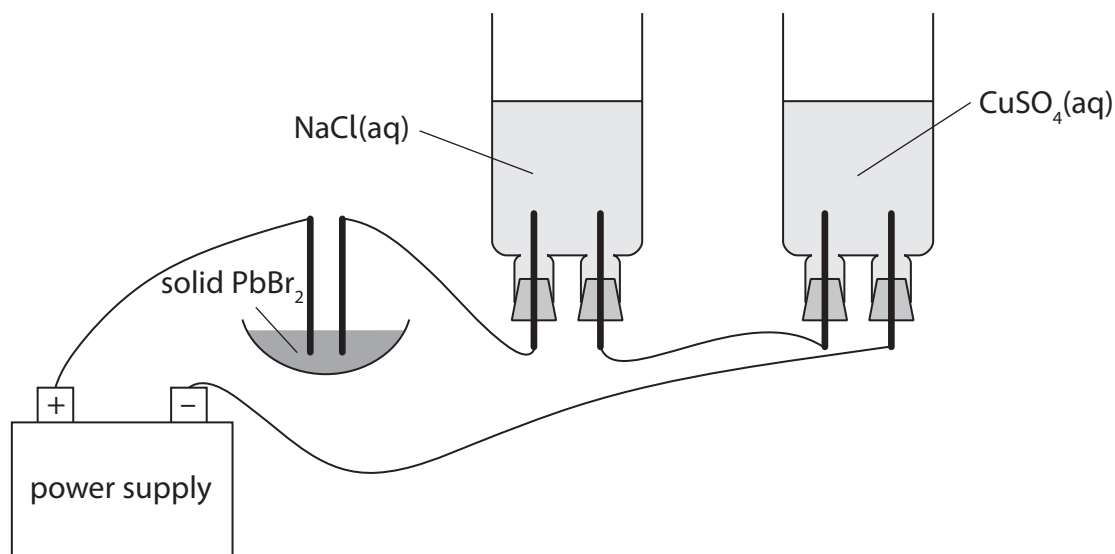
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(Total for Question 2 = 11 marks)



3 This apparatus can be used to investigate electrolysis.



(a) Name the particles that move through the connecting wires to form an electric current. (1)

(b) The electrodes are made of platinum, which is an inert metal.

State what is meant by the term **inert**.

(1)

(c) Explain why the electrolytic cell containing PbBr_2 needs to be heated before electrolysis can occur.

(2)

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(d) When NaCl(aq) is electrolysed, two gases form at the positive electrode and one gas forms at the negative electrode.

The formulae of the species in NaCl(aq) are Na⁺, Cl⁻, H⁺, OH⁻ and H₂O.

(i) Name the gases formed at each electrode. (2)

positive electrode and

negative electrode.....

(ii) Give ionic half-equations to show the formation of each gas. (3)

.....

.....

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.....

(e) The ionic half-equation for one of the reactions in the cell containing copper(II) sulfate solution is



During the electrolysis, a charge of 0.040 faradays passes through this cell.

Calculate the mass of copper metal formed. (2)

mass of copper = g

(Total for Question 3 = 11 marks)



- 4 Compounds containing C=C double bonds are used to manufacture alcohols such as ethanol and addition polymers such as PVC.

The table shows the formulae of some compounds containing C=C bonds.

A $\text{CH}_2=\text{CH}_2$	B $\begin{array}{c} \text{H} \quad \text{H} \quad \text{H} \\ \quad \quad \\ \text{H}-\text{C}=\text{C}-\text{C}-\text{H} \\ \\ \text{H} \end{array}$
C C_4H_8	D $\text{C}_2\text{H}_3\text{Cl}$

- (a) (i) Explain which three of these compounds are hydrocarbons.

(2)

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- (ii) Explain which compound is shown as a displayed formula.

(2)

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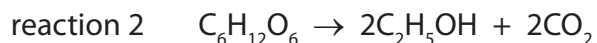
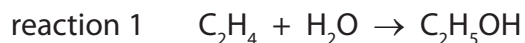
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- (b) Ethanol can be manufactured from compound A using reaction 1.
Ethanol can also be manufactured from glucose using reaction 2.

The equations for these reactions are



Give two advantages of using each reaction to manufacture ethanol.

(4)

reaction 1

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.....

.....

.....

reaction 2

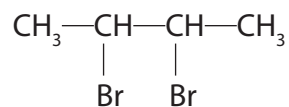
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- (c) Compound C reacts with bromine to form a product with this formula.



- (i) Use this formula to determine the name of compound C.

(2)

.....

- (ii) State the colour of the product formed when compound C reacts with bromine.

(1)

.....

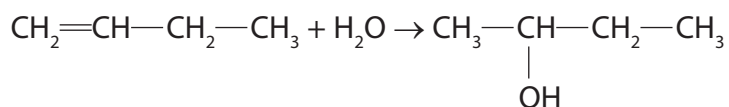


(d) Compound X is an isomer of compound C.

(i) Explain what is meant by the term **isomers**.

(2)

(ii) The equation for the conversion of compound X to an alcohol is



Place ticks (✓) in the boxes to show the two correct descriptions of this reaction.

(2)

addition	
dehydration	
hydration	
oxidation	
reduction	

(e) Compound D, chloroethene, can be used to manufacture the polymer PVC.

(i) State the full name of PVC.

(1)

(ii) Write an equation, containing displayed formulae, to represent the reaction that occurs in the manufacture of PVC.

(3)

(Total for Question 4 = 19 marks)



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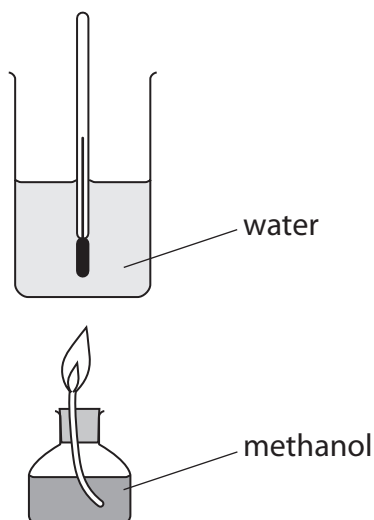
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P 4 8 0 8 3 A 0 1 3 2 0

- 5 A student uses this apparatus to find the increase in temperature of water when methanol, CH_3OH , is burned.



- (a) There are several reasons why the increase in temperature is less than expected.
- (i) One reason is the incomplete combustion of methanol to form only carbon monoxide and water.

Write the chemical equation for this incomplete combustion.

(2)

- (ii) State another reason why the increase in temperature is less than expected.

(1)

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(b) The student records these results.

mass of burner and methanol before combustion	84.7 g
mass of burner and methanol after combustion	83.2 g
mass of water	125 g
temperature of water at start	22 °C
temperature of water at end	58 °C

(i) Calculate the heat energy change (Q), in joules, in this experiment using the expression

$$Q = m \times 4.2 \times \Delta T$$

where m is the mass of water in grams and ΔT represents the increase in temperature.

(2)

$$Q = \dots\dots\dots \text{ J}$$

(ii) The relative molecular mass of methanol is 32

Use this information and your value for Q to calculate the molar enthalpy change, ΔH , for the combustion of methanol.

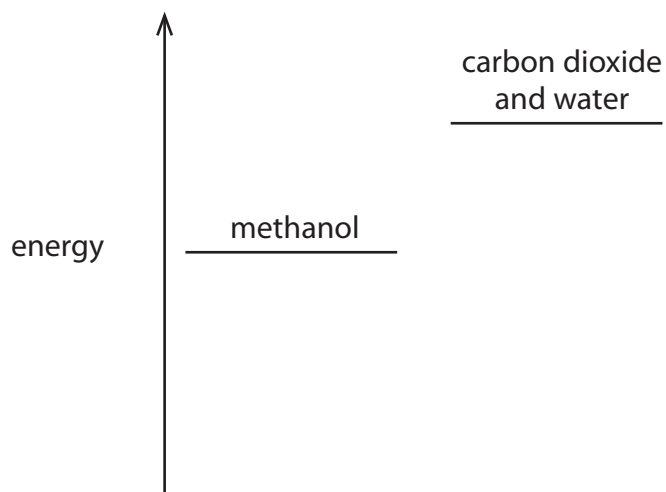
Give your answer in kJ/mol.

(4)

$$\Delta H = \dots\dots\dots \text{ kJ/mol}$$



(iii) The student draws an energy level diagram for the complete combustion of methanol.



Identify the two mistakes in his diagram.

(2)

1

2

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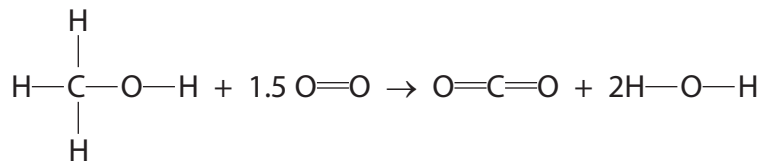
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(c) The student is given this table of average (mean) bond energies.

Bond	C—H	C—O	O—H	O=O	C=O
Average bond energy in kJ/mol	412	360	463	496	743

The equation for the complete combustion of methanol is



Use this equation and the information in the table to calculate another value for the molar enthalpy change, ΔH , for the combustion of methanol.

(4)

$\Delta H = \dots\dots\dots$ kJ/mol

(Total for Question 5 = 15 marks)

TOTAL FOR PAPER = 60 MARKS

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